

**FIRST SEMESTER CHEMISTRY FINAL EXAMINATION FOR GRADE 12<sup>TH</sup>****CHOOSE THE CORRECT ANSWERS**

- If the solute-solvent interactions are greater than the solute-solute and solvent-solvent interactions, what will be the total vapor pressure of the solution?
  - greater than that calculated from Raoult's law
  - less than that calculated from Raoult's law
  - The same as calculated from Raoult's law
  - Raoult's law cannot be applied for such interaction
- What is the freezing point of a solution of 250g of  $\text{CaCl}_2$  in 1.0kg of water ( $K_f$  for water =  $-1.86^\circ\text{C}/\text{kg}/\text{mol}$ )
  - $-1.3^\circ\text{C}$
  - $-6.5^\circ\text{C}$
  - $-9.0^\circ\text{C}$
  - $-12.55^\circ\text{C}$
- What is the molarity of a 5 g of hydrogen peroxide in 100mL solution that is used for hair bleaching
  - 0.015M
  - 0.15M
  - 1.5M
  - 3M
- What is the osmotic pressure, in atm, of a 0.30M solution of glucose in water that is used for intravenous infusion at body temperature,  $37^\circ\text{C}$ ?
  - 0.911 atm
  - 773.2atm
  - 7.6atm
  - 76atm
- Given the unbalanced equation  
 $\text{KMnO}_4 + \text{KI} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{I}_2 + \text{H}_2\text{O}$  How many grams of  $\text{KMnO}_4$  are needed to make 500mL of 0.250N solution?
  - 3.95g
  - 1.49g
  - 1.58g
  - 63.2g
- In which of the following cases will the dissolution of sugar be the most rapid?
  - Sugar crystals in hot water
  - Sugar crystals in cold water
  - Powdered sugar in hot water
  - Powdered sugar in cold water
- Which of the following is not a colligative property
  - Osmotic pressure
  - Vapour pressure
  - Boiling point
  - Freezing point depression
- What is the normality of a 1M solution of  $\text{H}_3\text{PO}_4$  given the following reaction  
 $\text{H}_3\text{PO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{HPO}_4 + \text{H}_2\text{O}$ 
  - 3N
  - 2N
  - 6N
  - 1N
- A given solution contains 100g of  $\text{NaCl}$  and 900g of water, what is the mole fraction of  $\text{NaCl}$ 
  - 0.0331
  - 0.9669
  - 0.3
  - 0.7
- A 0.029m aqueous solution of  $\text{K}_2\text{SO}_4$  has a freezing point of  $-0.14^\circ\text{C}$ , what are the value of  $i$  factors ( $K_{f\text{H}_2\text{O}} = -1.86^\circ\text{C}/\text{m}$ )
  - 2
  - 4
  - 2.3
  - 2.6
- Which of the following combination never form a buffer solution
  - $\text{HF}/\text{F}^-$
  - $\text{HCOOH}/\text{HCOO}^-$
  - $\text{HCO}_3^-/\text{CO}_3^{2-}$
  - $\text{HNO}_3/\text{NO}_3^-$
- Which one of the following are true about the graph of weak acid and strong base
  - The PH is equal to 7 at equivalence point
  - The graph indicates the titration of strong base with strong acid
  - The PH is greater than 7 at equivalence point
  - at half neutralization  $\text{PH} = \text{PK}_a$
- A solution in which propionic acid is 0.94% ionized has a PH of 2.85, what is the values of the acid ionization constant ( $K_a$ ) for propionic acid
  - $4.2 \times 10^{-5}$
  - $6.8 \times 10^{-4}$
  - $1.3 \times 10^{-5}$
  - $8.7 \times 10^{-5}$
- The acid-base indicator bromocresol green is a weak acid. The yellow acid and blue base forms of the indicator are present in equal concentrations in a solution when the PH is 4.68, what is the  $\text{PK}_a$  of bromocresol green
  - 4.48
  - 4.68
  - 5.68
  - 6.68
- For the step wise dissociation of aqueous  $\text{H}_3\text{PO}_4$ , which of the following is not conjugate acid-base pair
  - $\text{HPO}_4^{2-}$  and  $\text{H}_2\text{PO}_4^-$
  - $\text{H}_3\text{PO}_4$  and  $\text{H}_2\text{PO}_4^-$
  - $\text{H}_2\text{PO}_4^-$  and  $\text{HPO}_4^{2-}$
  - $\text{H}_2\text{PO}_4^-$  and  $\text{PO}_4^{3-}$
- To 0.2M solution of a weak monoprotic acid HA, enough quantity of its sodium salt NaA, was dissolved to give a concentration of 0.2M of the salt, what will be the acid concentration  $[\text{H}_3\text{O}^+]$  in the final solution ( $K_a$  of HA =  $1.8 \times 10^{-5}$ )
  - $1.8 \times 10^{-5}\text{M}$
  - $1.8 \times 10^{-4}\text{M}$
  - $2.8 \times 10^{-5}$
  - $2.8 \times 10^{-4}$
- Consider the following equilibria
  - $\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 + \text{OH}^-$
  - $\text{NH}_4^+ + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{NH}_3$
  - $\text{HSO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{H}_2\text{SO}_3$
 In which of the equilibrium does water acts as a Bronsted-lowry base?
  - I only
  - II only
  - II and III
  - I, II, III
- To a 50ml of a buffer solution that is 0.25M in HF and 0.1M in NaF, if 5ml of 0.05M NaOH is added, what is the final PH of this buffer solution,  $\text{PK}_a \text{ HF} = 3.17$ 
  - 2.77
  - 2.80
  - 2.74
  - 3.17
- When sodium acetate is added to an aqueous solution of acetic acid its PH
  - increases
  - Decreases
  - Remain constant
  - may increases or decreases
- How is a buffer solution prepared
  - By mixing a weak acid and strong acid
  - By mixing a weak acid and its conjugate base
  - By mixing a strong acid and its conjugate base
  - By mixing strong base and its conjugate acid
- Which one of the followings are not intensive properties
  - color
  - length
  - density
  - melting point

# 70 Question Chemistry Final Exam Answers

**Arturo Cuomo**



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