



Progress and Poverty

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 $2 \text{ ch}_4 + 2 \text{o}_2 \rightarrow 2 \text{co}_2 + 2 \text{h}_2\text{o}$
 $2 \text{ na} + \text{cl}_2 \rightarrow 2 \text{nacl}$
 $3 \text{ al} + \text{o}_2 \rightarrow \text{al}_2\text{o}_3$
 $4 \text{ al} + 3 \text{o}_2 \rightarrow 2 \text{al}_2\text{o}_3$
 $4 \text{ n}_2 + \text{h}_2 \rightarrow 2 \text{nh}_3$

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 $\text{h}_2 + \text{o}_2 \rightarrow \text{h}_2\text{o}$
 at the molecular level the reaction would look something like this notice that there are two oxygen atoms on the

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web in order to balance a chemical equation the quantities of each type of element and polyatomic ion that are present in the reactants and the products of the reaction must be determined the first reactant in this equation contains magnesium mg and a hydroxide ion oh⁻ which is a polyatomic anion

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 $\text{c}_3\text{h}_8 + \text{o}_2 \rightarrow 2 \text{h}_2\text{o} + 3 \text{co}_2$

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