

The following is a summary of most of the concepts, relationships, topics that you should know for this course. Note that this may include typographical errors because it has not been proof read and is a first draft, so please check all expressions with those in the book. I would only use this as a study guide and not as a replacement for studying the topic in depth within the book. Also note that I provide information about certain things I recommend you know how to do, calculate, derive, etc.  
*Prof. Musgrave*

#### Section 1.2a - Perfect gas law

Calculating relationships between  $p$ ,  $V$  and  $T$  at one state and another.

The molecules in perfect gases experience no interactions.

Do not apply the perfect gas law to systems where it does not apply (liquids, solids, dense gases)

#### Section 1.2b - Kinetic model of gases

Calculating average speeds of gas molecules as a function of  $M$  and  $T$ .

#### Section 1.2c - Mixtures of gases

Calculating partial pressures from total pressure and composition (mole fractions)

Dalton's Law: The sum of the pressures of gases of each component alone in a vessel sum to the total pressure of the mixture of those gases.

#### Section 1.3a - The compression factor

The volumes of real gases are different from perfect gases by the compression factor.

The compression factor is related to the Virial equation of state

#### Section 1.4a - van der Waals equation

Corrections for the size of molecules and their interactions can be included to form a more accurate description of the equation of state  $p(V,T)$

There is no closed form expression of  $V(p,T)$  - solve for  $V$  numerically.

There are closed forms of  $p(T,V)$  and  $T(p,V)$  so inverting partial derivatives is sometimes convenient to determine derivatives of  $V$  with respect to  $T$  or  $p$ .

#### Section 1.4b - van der Waals equation features

The vdW equation of state differs from the ideal gas law at high  $p$  and low  $T$  (where the assumption of ideal behavior breaks down).

### Chapter 1: Key Equations:

Perfect gas law:  $pV = nRT$

Temperature scales:  $T(K) = T(C) + 273.15$

Partial pressures:  $p = \sum p_i$  and  $p_i = x_i p = \frac{n_i}{n} p$

Van der Waals Equation of state:  $p = nRT / (V - nb) - an^2 / V^2$

#### Section 2.1 - Work, heat and energy

The energy of a system is its capacity to do work

Work is the transfer of energy by moving against a force - it is the transfer of energy by organized motion.

Heat is the transfer of random, thermal motion. It transfers disordered energy.

An adiabatic process transfers no heat and thus no random energy or entropy.

An isothermal process involves no change in  $T$ .

#### Section 2.2 - Internal energy $U$

Internal energy ( $U$ ) is the total energy of the system. It includes kinetic energy and potential energy and is a state of the system. It increases because either work is done on the system or heat is transferred to the system.

Section 2.2a - The internal energy of a perfect gas depends on  $T$  only -  $U(T)$ .

# Pchem Study Guide

**RJ Shavelson**



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## **Table of Contents Pchem Study Guide**

1. Understanding the eBook Pchem Study Guide
  - The Rise of Digital Reading Pchem Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Pchem Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pchem Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pchem Study Guide
  - Personalized Recommendations
  - Pchem Study Guide User Reviews and Ratings
  - Pchem Study Guide and Bestseller Lists
5. Accessing Pchem Study Guide Free and Paid eBooks

- Pchem Study Guide Public Domain eBooks
- Pchem Study Guide eBook Subscription Services
- Pchem Study Guide Budget-Friendly Options
- 6. Navigating Pchem Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Pchem Study Guide Compatibility with Devices
  - Pchem Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pchem Study Guide
  - Highlighting and Note-Taking Pchem Study Guide
  - Interactive Elements Pchem Study Guide
- 8. Staying Engaged with Pchem Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pchem Study Guide
- 9. Balancing eBooks and Physical Books Pchem Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pchem Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pchem Study Guide
  - Setting Reading Goals Pchem Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pchem Study Guide
  - Fact-Checking eBook Content of Pchem Study Guide
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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