

Rate of Reaction

The **rate** of a chemical reaction is how fast the reactants are changed into products.

Some reactions happen really quickly (e.g. an explosion). Other reactions happen slowly (e.g. the rusting of iron).

Collision Theory

Collision theory states that the rate of a reaction depends on two things.

1. The **collision frequency** of the reacting particles (how often they collide).
2. The **energy transferred** during a collision – particles have to collide with enough energy otherwise they won't react.

Factors Affecting the Rate of a Reaction

There are four main factors that affect the rate of a reaction.

1) Temperature

The higher the temperature, the faster the reaction. Increasing the temperature causes the particles to move more quickly and collide more often and with more energy.

2) Concentration

A higher concentration (or higher pressure for gases) of reactants causes the reaction to happen more quickly. This is because the reactant particles are more likely to collide if there are more of them.

3) Surface Area

The larger the surface area, the faster the reaction happens. Breaking up a solid reactant into very small pieces increases its surface area to volume ratio which increases the rate of the reaction.

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