



# CHEMISTRY CHEAT SHEET



# EQUILIBRIUM

## EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol  $K_c$ .

What is Equilibrium?		Equilibrium Constant $K_c$	Graphical Representation of Equilibrium
<p><b>What is Equilibrium?</b> Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol <math>K_c</math>.</p> <p><b>Equilibrium Constant <math>K_c</math></b> The equilibrium constant <math>K_c</math> is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p><b>Equilibrium Constant <math>K_c</math></b> The equilibrium constant <math>K_c</math> is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p><b>Equilibrium Constant <math>K_c</math></b> The equilibrium constant <math>K_c</math> is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p>	<p><b>Equilibrium Constant <math>K_c</math></b> The equilibrium constant <math>K_c</math> is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p><b>Equilibrium Constant <math>K_c</math></b> The equilibrium constant <math>K_c</math> is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p><b>Equilibrium Constant <math>K_c</math></b> The equilibrium constant <math>K_c</math> is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p>	<p><b>Equilibrium Constant <math>K_c</math></b> The equilibrium constant <math>K_c</math> is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p><b>Equilibrium Constant <math>K_c</math></b> The equilibrium constant <math>K_c</math> is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p><b>Equilibrium Constant <math>K_c</math></b> The equilibrium constant <math>K_c</math> is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p>	
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Factor	Effect on Equilibrium	Example
Concentration	<ul style="list-style-type: none"> <li>Increasing the concentration of a reactant shifts the equilibrium to the right.</li> <li>Increasing the concentration of a product shifts the equilibrium to the left.</li> </ul>	$\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$ <p>Increasing the concentration of <math>\text{N}_2</math> shifts the equilibrium to the right, producing more <math>\text{NH}_3</math>.</p>
Pressure	<ul style="list-style-type: none"> <li>Increasing the pressure shifts the equilibrium to the side with fewer moles of gas.</li> <li>Decreasing the pressure shifts the equilibrium to the side with more moles of gas.</li> </ul>	$\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$ <p>Increasing the pressure shifts the equilibrium to the right, producing more <math>\text{NH}_3</math>.</p>
Temperature	<ul style="list-style-type: none"> <li>Increasing the temperature shifts the equilibrium to the endothermic side.</li> <li>Decreasing the temperature shifts the equilibrium to the exothermic side.</li> </ul>	$\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$ <p>Increasing the temperature shifts the equilibrium to the left, producing less <math>\text{NH}_3</math>.</p>

# Modern Chemistry Study Guide Chemical Equilibrium

**Padhraic Smyth**



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