

## ANSWERS

### Nuclear Chemistry Review Questions

1. Compare the mass, charge, and penetration power of alpha, beta and gamma radiation.

	Mass	Charge	Penetration
Alpha	4	+	Lowest
Beta	0	-	Middle
Gamma	0	0	Highest

2. In 5.49 seconds, 1.20g of argon-35 decays to leave only 0.15g. What is the half-life?

1.20g  $\rightarrow$  0.60g  $\rightarrow$  0.30g  $\rightarrow$  0.15g, therefore it took 3 half-life periods.

$$5.49 \div 3 = 1.83 \text{ sec}$$

3. How many days does it take 16.0g of gold-198 to decay to 1.0g?

16.0g  $\rightarrow$  8.0g  $\rightarrow$  4.0g  $\rightarrow$  2.0g  $\rightarrow$  1.0g, therefore 4 half-life periods

From Ref Table N, 1 half-life = 2.69d

$$2.69\text{d} \times 4 = 10.76 \text{ days}$$

4. Why does an atom undergo radioactive decay?

Because it has an unstable nucleus

5. Compare and contrast fission and fusion. Why is fission used for electrical and fusion is not?

Fission=Division, the splitting of large unstable atoms(nucleus)

Fusion=joining together, small isotopes of H, creates way more energy

6. How is nuclear chemistry applied to medicine? What type of radioisotope must be used when it is injected into a person?

Radioisotopes are used to treat and diagnose illnesses

If injected they must have short half-lives and quickly removed from the body

7. Write out the nuclear decay of radium-226.



8. What is different about nuclear reactions when compared to a chemical reaction?

Nuclear reactions transform atoms into other atoms, chemical reactions involved the rearrangement of electrons

9. At what point do all isotopes of an element become radioactive?

Up to and including atomic # 83

# Nuclear Chemistry Answers

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**The Handy Chemistry Answer Book** Justin P. Lomont, Ian C. Stewart, 2013-10-01 Don't be mixed up about chemistry. Simplify the complex chemical reactions that take place everywhere in our lives with this engaging easy-to-follow question and answer guide. Where would we be without atoms and compounds? Gas, liquids, solids, and plasma. Acids and bases. Bonds and reactions. Matter and energy. The Handy Chemistry Answer Book covers the building blocks of life and the universe. The secret life of atoms, how polar bears aren't actually white, why oil and water don't mix, and much, much more are revealed and explained. This informative guide covers the basics of chemistry: history, atomic structures, chemical bonds, and reactions; organic and inorganic chemistry; to more advanced material: nuclear chemistry, biochemistry, physical and theoretical chemistry. By answering nearly 1,000 common chemistry questions including: What causes lightning? How does photosynthesis work? What are hard and soft Lewis acids and bases? What makes a fabric waterproof? What are the twelve principles of green chemistry? When did alchemists finally abandon trying to make gold? What is Le Chatelier's principle? What do the different octane ratings mean at the gas pump? What is genetic engineering? Why is calcium important for strong bones? What is the 18-electron rule? Why does chocolate turn white as it ages? Chemical reactions that rule the world: their properties, structure, composition, behavior, and history are tackled and explained in plain English in The Handy Chemistry Answer Book. With many photos, illustrations, a few formulas, molecular diagrams, and other graphics, this fun fact-filled tome is richly illustrated. A history of chemistry, timeline, appendices on Nobel Prize winners, a bibliography, further reading section, glossary of terms, a table of physical constants, a table of conversion factors, and an extensive index add to its usefulness.

**Radiochemistry and Nuclear Chemistry** Gregory Choppin, Jan-Olov Liljenzin, Jan Rydberg, 2016-01-26 Nuclear chemistry comprises isotope chemistry, radiochemistry, radiation chemistry, and nuclear reaction chemistry, along with applications. These interrelated fields are all covered in this textbook for chemists and chemical engineers. This new edition of the standard work Nuclear Chemistry has been completely rewritten and restructured to suit teaching and learning needs in a wide range of chemistry courses, such as basic courses in radiochemistry or more advanced nuclear chemistry courses. The book is divided into sections that closely fit teaching demands. The first chapter gives a broad introduction and background to the subject, and the second chapter covers stable isotopes. Chapters 3 to 9 comprise what is generally regarded as radiochemistry. Chapters 10 to 17 offer a course in nuclear reaction chemistry. Chapter 18 deals with biological radiation effects for the chemist. The last four chapters give a guide to nuclear energy: energy production, fuel cycle, waste management, the largest applied field of nuclear chemistry. Over 200 exercises with model answers remain largely unchanged from the first edition, so teachers working from the earlier text should find only advantages in switching to this new restructured coursebook on all aspects of nuclear chemistry. The book fully meets the authors' objectives: it is well written in a logical, objective, thought-provoking, and quite easily readable style. It should appeal to the serious student of radio and

nuclear chemistry at either undergraduate or postgraduate level as well as to readers with a more general interest in nuclear science and its impact on the environment Applied Radiation and Isotopes July 1995 This book is an excellent readable account of a significant part of the scientific achievements of more than half this century The authors have dedicated the book to Nobel Laureate Glenn T Seaborg and its scholarship makes it a fitting tribute Radiological Protection Bulletin December 1995

**Chemistry** John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

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