

Chapter 12 Study Guide

1) A, $d=5$

2) N

3) N

4) A, $d=-4$

5) G, $r=1/2$

6) N

7) A, $d=2$

8) N

9) N

10) A $d = \frac{2}{3}$

Notebook Guide Answers

M Walker



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web the pcl₃ lewis structure has the typical case of phosphorus p in the center with 3 bonds to 3 other atoms phosphorus is
from the same column as nitrogen in the periodic table meaning that p and n generally have the same bonding structure note
the lone pair dots without bonds on top of p just like for n in the previous example for nh₃ chlorine cl

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web step 1 figure out how many electrons the molecule must have based on the number of valence electrons in each atom
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