

FIRST SEMESTER CHEMISTRY FINAL EXAMINATION FOR GRADE 12TH**CHOOSE THE CORRECT ANSWERS**

- If the solute-solvent interactions are greater than the solute-solute and solvent-solvent interactions, what will be the total vapor pressure of the solution?
A. greater than that calculated from Raoult's law B. less than that calculated from Raoult's law
C. The same as calculated from Raoult's law D. Raoult's law cannot be applied for such interaction
- What is the freezing point of a solution of 250g of CaCl_2 in 1.0kg of water (K_f for water = $-1.86^\circ\text{C}/\text{kg}/\text{mol}$) A. -1.3°C B. -6.5°C C. -9.0°C D. -12.55°C
- What is the molarity of a 5 g of hydrogen peroxide in 100mL solution that is used for hair bleaching A. 0.015M B. 0.15M C. 1.5M D. 3M
- What is the osmotic pressure, in atm, of a 0.30M solution of glucose in water that is used for intravenous infusion at body temperature, 37°C ?
A. 0.911 atm B. 773.2atm C. 7.6atm D. 76atm
- Given the unbalanced equation
 $\text{KMnO}_4 + \text{KI} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{I}_2 + \text{H}_2\text{O}$ How many grams of KMnO_4 are needed to make 500mL of 0.250N solution? A. 3.95g B. 1.49g C. 1.58g D. 63.2g
- In which of the following cases will the dissolution of sugar be the most rapid?
A. Sugar crystals in hot water B. Sugar crystals in cold water
C. Powdered sugar in hot water D. Powdered sugar in cold water
- Which of the following is not a colligative property
A. Osmotic pressure B. Vapour pressure C. Boiling point D. Freezing point depression
- What is the normality of a 1M solution of H_3PO_4 given the following reaction
 $\text{H}_3\text{PO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{HPO}_4 + \text{H}_2\text{O}$ A. 3N B. 2N C. 6N D. 1N
- A given solution contains 100g of NaCl and 900g of water. what is the mole fraction of NaCl
A. 0.0331 B. 0.9669 C. 0.3 D. 0.7
- A 0.029m aqueous solution of K_2SO_4 has a freezing point of -0.14°C . what are the value of i factors ($K_{f,\text{H}_2\text{O}} = -1.86^\circ\text{C}/\text{m}$) A. 2 B. 4 C. 2.3 D. 2.6
- Which of the following combination never form a buffer solution
A. HF/F^- B. $\text{HCOOH}/\text{HCOO}^-$ C. $\text{HCO}_3^-/\text{CO}_3^{2-}$ D. $\text{HNO}_3/\text{NO}_3^-$
- Which one of the following are true about the graph of weak acid and strong base
A. The PH is equal to 7 at equivalence point B. The graph indicates the titration of strong base with strong acid
C. The PH is greater than 7 at equivalence point D. at half neutralization $\text{PH} = \text{PK}_a$
- A solution in which propionic acid is 0.94% ionized has a PH of 2.85. what is the values of the acid ionization constant (K_a) for propionic acid A. 4.2×10^{-4} B. 6.8×10^{-4} C. 1.3×10^{-5} D. 8.7×10^{-5}
- The acid-base indicator bromocresol green is a weak acid. The yellow acid and blue base forms of the indicator are present in equal concentrations in a solution when the PH is 4.68. what is the PK_a of bromocresol green A. 4.48 B. 4.68 C. 5.68 D. 6.68
- For the step wise dissociation of aqueous H_3PO_4 , which of the following is not conjugate acid-base pair A. HPO_4^{2-} and H_2PO_4^- B. H_3PO_4 and H_2PO_4^- C. H_2PO_4^- and HPO_4^{2-} D. H_2PO_4^- and PO_4^{3-}
- To 0.2M solution of a weak monoprotic acid HA, enough quantity of its sodium salt NaA, was dissolved to give a concentration of 0.2M of the salt. what will be the acid concentration $[\text{H}_3\text{O}^+]$ in the final solution (K_a of HA = 1.8×10^{-5})
A. $1.8 \times 10^{-5}\text{M}$ B. $1.8 \times 10^{-4}\text{M}$ C. 2.8×10^{-5} D. 2.8×10^{-4}
0.0057M B. 0.057M C. 0.57M D. 5.70M
- Consider the following equilibria I. $\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 + \text{OH}^-$ II. $\text{NH}_4^+ + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{NH}_3$,
III. $\text{HSO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{H}_2\text{SO}_3$ In which of the equilibrium does water acts as a Bronsted-lowry base?
A. I only B. II only C. II and III D. I, II, III
- To a 50ml of a buffer solution that is 0.25M in HF and 0.1M in NaF, if 5ml of 0.05M NaOH is added, what is the final PH of this buffer solution, $\text{PK}_a \text{ HF} = 3.17$ A. 2.77 B. 2.80 C. 2.74 D. 3.17
- When sodium acetate is added to an aqueous solution of acetic acid its PH
A. increases B. Decreases C. Remain constant D. may increases or decreases
- How is a buffer solution prepared A. By mixing a weak acid and strong acid B. By mixing a weak acid and its conjugate base C. By mixing a strong acid and its conjugate base D. By mixing strong base and its conjugate acid
- Which one of the followings are not intensive properties A. color B. length C. density D. melting point

70 Question Chemistry Final Exam Answers

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Encyclopedia of Chemical Processing and Design John J. McKetta Jr,1993-06-02 Written by engineers for engineers with over 150 International Editorial Advisory Board members this highly lauded resource provides up to the minute information on the chemical processes methods practices products and standards in the chemical and related industries

Report Upon the Survey of the University of Wisconsin Wisconsin. State Board of Public Affairs,1914 114
Topic-wise Quantitative Aptitude & Data Interpretation Previous Year Question Bank for IBPS/ SBI/ RRB/ RBI Bank Clerk/ PO Prelim & Main Exams (2010 - 2024) 8th Edition | 100% Solved Quant & DI PYQs Disha Experts,2024-05-02 The thoroughly revised updated 8th edition 114 Quantitative Aptitude Data Interpretation Topic wise Previous Year Solved Papers for IBPS SBI Bank PO Clerk Prelim Main Exams 2010 23 consists of past solved papers for Prelim and Main Exams of Banks IBPS PO IBPS Clerk SBI PO SBI Clerk IBPS RRB PO IBPS RRB Office Assistant RBI Assistant and IBPS SBI Specialist Officer from 2010 to 2024 The coverage of the papers has been kept RECENT 2010 to 2024 as they actually reflect the changed pattern of the Banking exams Thus the papers prior to 2010 have not been included in the book In all there are 114 Question Papers having 4700 Questions from 2010 to 2024 which have been divided into 15 Topics with detailed solutions Practicing these questions aspirants will come to know about the pattern and toughness of the questions asked in the bank examinations In the end this book will make the aspirants competent enough to crack the these Entrance Examination with good score The strength of the book lies in the originality of its question papers and Errorless Solutions The solution of each and every question is provided in detail step by step so as to provide 100% concept clarity to the students *State Board Journal of America* ,1907 *Cracking the DAT (Dental Admission Test)* Princeton Review (Firm),2012 Includes perforated quick reference study sheets The Hidden Curriculum - Faculty Made Tests in Science Sheila Tobias,1997 **Quantitative Aptitude & Data Interpretation Topic-wise Solved Papers for IBPS/ SBI Bank PO/ Clerk Prelim & Main Exams (2010-20) 4th Edition** Disha Experts,2020-04-18 **5 Steps to a 5: AP Chemistry 2021** Mary Millhollon,Richard H. Langley,2020-10-02 MATCHES THE LATEST EXAM In this hybrid year let us supplement your AP classroom experience with this multi platform study guide The immensely popular 5 Steps to a 5 AP Chemistry guide has been updated for the 2020 21 school year and now contains 3 full length practice exams available both in the book and online that reflect the latest exam Up to Date Resources for COVID 19 Exam Disruption Access to a robust online platform Comprehensive overview of the AP Chemistry exam format Hundreds of practice exercises with thorough answer explanations Proven strategies specific to each section of the test A self guided study plan including flashcards games and more online **Kaplan GRE Exam 2008 Comprehensive Program** Kaplan,2007-07-03 Students who turn to Kaplan can depend on the most accurate information on the GRE test change As always Kaplans comprehensive guides combine proven score raising strategies with the latest technology to help students exceed their expectations **The American Bookseller** ,1886 **Educational**

Measurements and the Classroom Teacher Adam Raymond Gilliland, Riverda Harding Jordan, 1924 *Catalogue of the School Bulletin Publications ...* Charles William Bardeen, 1900 *Midland Druggist and the Pharmaceutical Review*, 1923 *Chemical Engineering and Mining Review*, 1924 **The Publishers Weekly**, 1924 Journal of the American Medical Association American Medical Association, 1908 Includes proceedings of the Association papers read at the annual sessions and list of current medical literature 5 Steps to a 5 AP Chemistry John Moore, 2003-08-22 For the more than one million students taking the AP exams each year Boxed quotes offering advice from students who have aced the exams and from AP teachers and college professors Sample tests that closely simulate real exams Review material based on the contents of the most recent tests Icons highlighting important facts vocabulary and frequently asked questions Websites and links to valuable online test resources along with author e mail addresses for students with follow up questions Authors who are either AP course instructors or exam developers **The Popular Science Monthly**, 1923 Anesthesiology Examination and Board Review 7/E Mark Dershwitz, J. Matthias Walz, 2014-02-05 The Single Best Review for the American Board of Anesthesiology Written Examinations Completely Updated to Reflect the New Exam Format Anesthesiology Examination and Board Review is the most time proven and effective way to prepare for the ABA written examinations This powerful study guide delivers 1 500 board style multiple choice questions with referenced paragraph length discussions for each answer Anesthesiology Examination and Board Review has been completely redesigned and rewritten to reflect the new exam format Like the exam this edition is divided into two parts Basic and Advanced each concluding with a practice test There is no better way for you to boost your test taking skills subject knowledge and confidence than this essential study partner FEATURES 1 500 board style questions with fully explained referenced answers New information from the American Board of Anesthesiology on changes to the certification and recertification process and requirements Complete coverage of the entire field of anesthesiology Two comprehensive practice tests that mimic the actual exams and allow you to test your skill at answering questions under simulated exam conditions Valuable test taking strategies New questions that reflect the latest research and advances **Colliery Engineer**, 1894

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Table of Contents 70 Question Chemistry Final Exam Answers

1. Understanding the eBook 70 Question Chemistry Final Exam Answers
 - The Rise of Digital Reading 70 Question Chemistry Final Exam Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 70 Question Chemistry Final Exam Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 70 Question Chemistry Final Exam Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 70 Question Chemistry Final Exam Answers
 - Personalized Recommendations
 - 70 Question Chemistry Final Exam Answers User Reviews and Ratings
 - 70 Question Chemistry Final Exam Answers and Bestseller Lists
5. Accessing 70 Question Chemistry Final Exam Answers Free and Paid eBooks
 - 70 Question Chemistry Final Exam Answers Public Domain eBooks
 - 70 Question Chemistry Final Exam Answers eBook Subscription Services
 - 70 Question Chemistry Final Exam Answers Budget-Friendly Options

6. Navigating 70 Question Chemistry Final Exam Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 70 Question Chemistry Final Exam Answers Compatibility with Devices
 - 70 Question Chemistry Final Exam Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 70 Question Chemistry Final Exam Answers
 - Highlighting and Note-Taking 70 Question Chemistry Final Exam Answers
 - Interactive Elements 70 Question Chemistry Final Exam Answers
8. Staying Engaged with 70 Question Chemistry Final Exam Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 70 Question Chemistry Final Exam Answers
9. Balancing eBooks and Physical Books 70 Question Chemistry Final Exam Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 70 Question Chemistry Final Exam Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 70 Question Chemistry Final Exam Answers
 - Setting Reading Goals 70 Question Chemistry Final Exam Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 70 Question Chemistry Final Exam Answers
 - Fact-Checking eBook Content of 70 Question Chemistry Final Exam Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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