

Chapter 7

Redox Reaction

Question 7.1:

Assign oxidation number to the underlined elements in each of the following species:

- (a) $\text{NaH}_2\underline{\text{P}}\text{O}_4$ (b) $\text{NaH}\underline{\text{S}}\text{O}_4$ (c) $\text{H}_4\underline{\text{P}}_2\text{O}_7$ (d) $\text{K}_2\underline{\text{Mn}}\text{O}_4$
(e) $\text{Ca}\underline{\text{O}}_2$ (f) $\text{Na}\underline{\text{B}}\text{H}_4$ (g) $\text{H}_2\underline{\text{S}}_2\text{O}_7$ (h) $\text{KAl}(\underline{\text{S}}\text{O}_4)_2 \cdot 12 \text{H}_2\text{O}$

Answer 7.1:

- (a) $\text{NaH}_2\underline{\text{P}}\text{O}_4$

Let the oxidation number of P be x.

We know that,

Oxidation number of Na = +1

Oxidation number of H = +1

Oxidation number of O = -2



Then, we have

$$1(+1) + 2(+1) + 1(x) + 4(-2) = 0 \Rightarrow 1 + 2 + x - 8 = 0 \Rightarrow x = +5$$

Hence, the oxidation number of P is +5.

- (b) $\text{NaH}\underline{\text{S}}\text{O}_4$



Then, we have

$$1(+1) + 1(+1) + 1(x) + 4(-2) = 0 \Rightarrow 1 + 1 + x - 8 = 0 \Rightarrow x = +6$$

Hence, the oxidation number of S is +6.

- (c) $\text{H}_4\underline{\text{P}}_2\text{O}_7$



Then, we have

$$4(+1) + 2(x) + 7(-2) = 0$$

$$\Rightarrow 4 + 2x - 14 = 0 \Rightarrow 2x = +10 \Rightarrow x = +5$$

Hence, the oxidation number of P is +5.

- (d) $\text{K}_2\underline{\text{Mn}}\text{O}_4$



Then, we have

$$2(+1) + x + 4(-2) = 0 \Rightarrow 2 + x - 8 = 0 \Rightarrow x = +6$$

Hence, the oxidation number of Mn is +6.

Redox Reaction Solutions Manual

Susan McMurry



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structure of chemical substances and the changes both chemical and physical that they undergo

Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light

Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth

Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O

Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms

Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

Student Solutions Manual: Ssm Chemistry Deborah Wiegand, 2003 This manual contains complete worked out solutions to all follow up problems and about half of all the chapter problems Each chapter of solutions opens with a summary of the text chapter content and a list of key equations needed to solve the problems

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