

PERIODIC TABLE BASIC

Reading the Periodic Table:

Objectives:

1. Explain why elements in a group have similar properties.
2. Identify the four blocks of the periodic table.

The Periodic Table:

Our current periodic table lists 109 elements. It is arranged according to the periodic law:

- Elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern.
- Elements that have similar properties are arranged in groups or families - vertical columns.
- The horizontal rows are called periods.
- Each period contains more elements than the previous one. This has to do with the increasing energy levels ($n=1, \dots$) of the atom.
- period 1 - 2 elements
- period 2 & 3 - 8 elements
- period 4 & 5 - 18 elements
- period 6 - 32 elements

Labeling and Naming:

There are currently 3 numbering schemes used in chemistry for the ordering of groups.

- Roman - IA, VIIIA, & IB-VIIB
- Arabic - 1A-2A, 3B-8B, 1B-2B, 3A-8A
- Number - 1-18

Families:

- 1A - alkali metals
- 2A - alkaline earth metals
- 7A - halogens
- 8A - noble gases.
- 4A - sometimes known as the Carbon group

Exception - Hydrogen can be better placed in group 7A. Although, some books place it in both.

Metals & Non-metals:

Metals include the areas shaded light blue and include the families 1A & 2A. They share many characteristics.

- luster (shine)
- good conductors of heat and electricity
- typically solid at room temperature
- most are malleable and ductile

Non-metals are shaded light red in most periodic tables and form a diagonal from C-6 to Rn-86 and include all the elements of that corner including H. They share some of these characteristics.

- do not possess luster
- poor conductors of heat and electricity
- neither malleable or ductile
- many are gasses at room temperature
- can be solid, semisolid, gas, or liquid (Br). One is the hardest material on Earth

Semimetals or metalloids are shaded light purple in most periodic tables form a diagonal from B-5 to At-85 and include all the elements Ge-32 and Sb-51. They share some of these characteristics.

have some properties of metals and some of non-metals

Electron Configuration:

As you learned earlier, the number of protons in an atom is equal to the number of electrons in the atom if it is neutral. The periodic table allows you to predict the bonding of an atom by allowing you to see the number of electrons in the valence orbital. The valence orbital is the highest, outermost orbital of the atom and is the only one involved in the bonding of the atom. Here are some hints for predicting valence electrons:

- Elements of the same group (column) have the same valence electrons or configuration
- s-block is group 1A & 2A
- p-block is group 3A to 8A
- These are the representative elements or main group elements
- d-block is group 3B to 2B (look at the periodic table above because the numbers are not sequential)
- These are the transition metals
- f-block is located below the main table
- These are the inner transition metals or rare earth elements

To save space on electron configuration notation you can use the abbreviated notation by listing the inner gas core and then the valence electrons.

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Table of Contents Periodic Table Basic Speedy Study Guide

1. Understanding the eBook Periodic Table Basic Speedy Study Guide
 - The Rise of Digital Reading Periodic Table Basic Speedy Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Basic Speedy Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Basic Speedy Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Basic Speedy Study Guide
 - Personalized Recommendations
 - Periodic Table Basic Speedy Study Guide User Reviews and Ratings
 - Periodic Table Basic Speedy Study Guide and Bestseller Lists
5. Accessing Periodic Table Basic Speedy Study Guide Free and Paid eBooks
 - Periodic Table Basic Speedy Study Guide Public Domain eBooks
 - Periodic Table Basic Speedy Study Guide eBook Subscription Services
 - Periodic Table Basic Speedy Study Guide Budget-Friendly Options

6. Navigating Periodic Table Basic Speedy Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Basic Speedy Study Guide Compatibility with Devices
 - Periodic Table Basic Speedy Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Basic Speedy Study Guide
 - Highlighting and Note-Taking Periodic Table Basic Speedy Study Guide
 - Interactive Elements Periodic Table Basic Speedy Study Guide
8. Staying Engaged with Periodic Table Basic Speedy Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Basic Speedy Study Guide
9. Balancing eBooks and Physical Books Periodic Table Basic Speedy Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Basic Speedy Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Basic Speedy Study Guide
 - Setting Reading Goals Periodic Table Basic Speedy Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Basic Speedy Study Guide
 - Fact-Checking eBook Content of Periodic Table Basic Speedy Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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