

## Chemistry 2<sup>nd</sup> Semester Final Review **KEY**

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

*Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)*

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)  
V1=3.8 L; P1=2.71 atm  
V2=1.47L; P2=? Atm  
Since all units are standard (L, atm), we do not need to convert the units before proceeding.  
 $(P_1V_1)/T_1=(P_2V_2)/T_2$   
T goes away (not mentioned in problem)  
 $P_1V_1=P_2V_2$   
Isolate for P2:  $P_2=(P_1V_1)/V_2$   
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})= 7.0 \text{ atm}$  (L cancelled and it's 2 sig figs)
2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)  
V1= 1.49 L, T1= 34.75 degrees Celsius  
V2= ? L, T2= 78.41 degrees Celsius  
We need to convert from Celsius to Kelvins before proceeding.  
T1= 34.75 + 273= 307.75 K  
T2= 78.41 + 273= 351.41 K  
 $(P_1V_1)/T_1=(P_2V_2)/T_2$   
Get rid of P  
 $V_1/T_1= V_2/T_2$   
Isolate V2:  $V_2=(V_1/T_1)T_2$   
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})= 1.70 \text{ L}$
3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?  
The combined gas law won't work here. You only have one condition and nothing changes.  
PV=nRT will work. (n= moles and R= 0.08206 L atm/K mol)  
P= 1.04 atm, V= ? L, n= OOPS we need to convert from grams to moles!, T= 84.7 degrees Celsius + 273= 357.7 K  
n: Hydrogen gas is diatomic!  
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/ 2.02 \text{ g H}_2)= 4.193069307 \text{ mol H}_2$   
Isolate V:  $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/ K mol} \times 357.7 \text{ K})/ 1.04 \text{ atm}= 118 \text{ L}$
4. What volume is occupied by 56.75 g of oxygen gas at STP?

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