

Reactions of Amines and Epoxides

Preparation of Amines

Amines of different types can be prepared by various methods. The following are some of the common methods:



Free electron lone-pairs

There are two types of bonding: **sigma bonding** and **covalent bonding**.

- **Sigma bonds** are formed by the overlap of orbitals from one atom to another (conspicuous) (green) and **covalent bonding** (purple). Example:



- **Covalent bonds** are formed by sharing of electrons between two atoms.



Sigma (σ) and pi (π) bonds

Covalent bonds are classified as **sigma (σ) bonds** and **pi (π) bonds**.

- **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **One single bond** is a sigma bond.
- **Two double bonds** consist of one sigma and one pi bond(s).



• pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be **polar** and **nonpolar**.



• If the difference in electronegativity is **less than 0.4**, the electrons are shared equally (distribution) for two atoms, forming a **nonpolar covalent bond**. (Not a typical example, see the C-C and C-H bonds).

• If the difference in electronegativity is **between 0.4 and 1.7**, the electrons are shared unequally, forming a **polar covalent bond**. (Not a typical example, see atoms for carbon a bonded to N, O, F and other halogens).

• A difference of **1.7 or higher** is so high that the electrons are not longer shared and a **ionic bond** is formed.

A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is described by the **dipole moment (μ)**.

Let's consider a bond between two atoms:



Electrons are pulled toward the more electronegative atom.

• The dipole moment is shown as a vector pointing from the less electronegative to toward the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).

• δ⁻ (delta minus) indicates an electron-deficient area.

• δ⁺ (delta plus) indicates an electron-rich area.

Molecular Dipole

To increase the molecular polarity, the same side of the dipole is used.

The polarity of the molecule is supported. **Geometrical molecules** are often a polar case for dipole moment are considered. A good example is carbon dioxide.



Lower the value of the dipole moment indicates the more non-polar.

Average values of C-C and C-H bonds, depending on hybridization			
C-C Bond	sp ³ -C-C	sp ² -C-C	sp-C-C
Hybridization	sp ³	sp ²	sp
Length	154 Å	147 Å	138 Å
Strength	376 kJ/mol	473 kJ/mol	505 kJ/mol
Hybridization	sp ³	sp ²	sp
C-H Bond	sp ³ -C-H	sp ² -C-H	sp-C-H
Hybridization	sp ³	sp ²	sp
Length	109 Å	108 Å	106 Å
Strength	413 kJ/mol	460 kJ/mol	505 kJ/mol

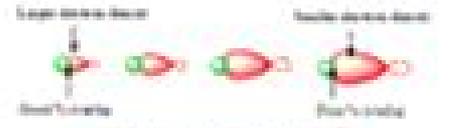
Higher electronegativity values and stronger bond →

Bond Length and Bond Strength

← Bond length increases as →

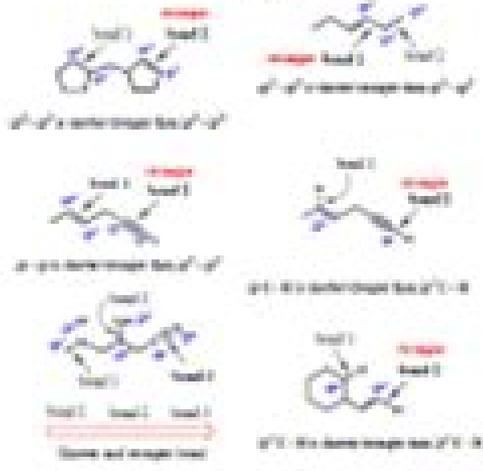
Bond	H-H	H-Cl	H-Br	H-I
Length	74.7 Å	127.6 Å	141.4 Å	154.9 Å
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds and stronger than triple bonds.



→ Stronger bond to weaker bond →

Example



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Brendan G. Carr



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