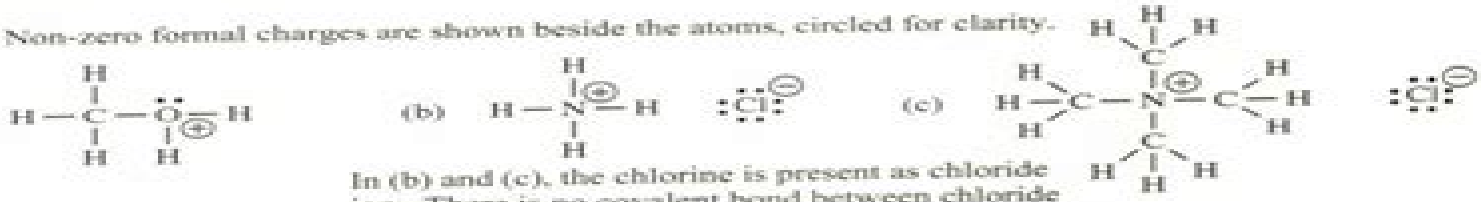
1-5 The symbols "δ⁺" and "δ⁻" indicate bond polarity by showing partial charge. (In the arrow symbolism, the arrow should point to the partial negative charge.)

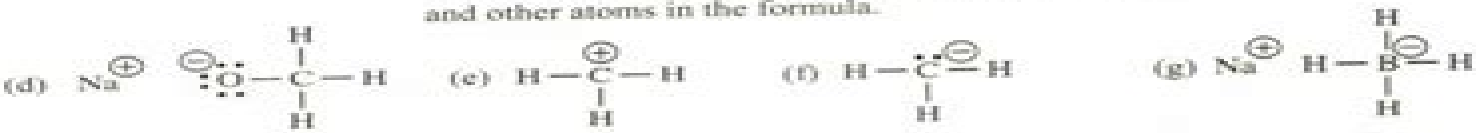
(a) $\overset{\delta+}{C}-\overset{\delta-}{Cl}$ (b) $\overset{\delta+}{C}-\overset{\delta-}{O}$ (c) $\overset{\delta+}{C}-\overset{\delta-}{N}$ (d) $\overset{\delta+}{C}-\overset{\delta-}{S}$ (e) $\overset{\delta-}{C}-\overset{\delta+}{B}$

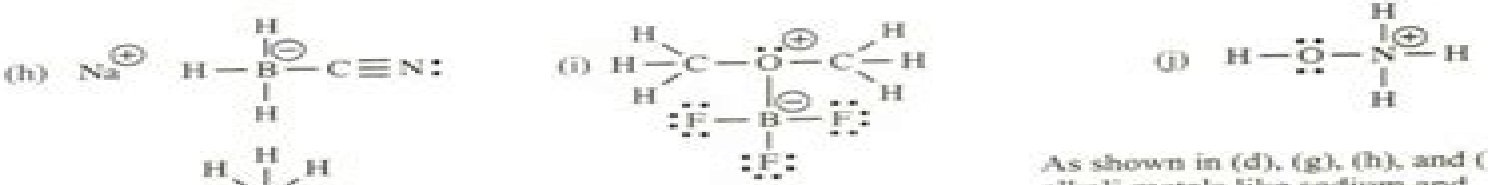
(f) $\overset{\delta+}{N}-\overset{\delta-}{Cl}$ (g) $\overset{\delta+}{N}-\overset{\delta-}{O}$ (h) $\overset{\delta-}{N}-\overset{\delta+}{S}$ (i) $\overset{\delta-}{N}-\overset{\delta+}{B}$ (j) $\overset{\delta+}{B}-\overset{\delta-}{Cl}$

1-6 Non-zero formal charges are shown beside the atoms, circled for clarity.



In (b) and (c), the chlorine is present as chloride ion. There is no covalent bond between chloride and other atoms in the formula.





As shown in (d), (g), (h), and (k), alkali metals like sodium and potassium form only ionic bonds, never covalent bonds.

1-7 Resonance forms in which all atoms have full octets are the most significant contributors. In resonance forms, ALL ATOMS KEEP THEIR POSITIONS—ONLY ELECTRONS ARE SHOWN IN DIFFERENT POSITIONS. (In this Solutions Manual, braces { } are commonly used to denote resonance forms.)

One more helpful hint when drawing resonance forms of structures with charges: *move electrons toward (+) charges and away from (−) charges.* Arrows show how to alter one structure to make it into the next one. In Chapter 1, arrows in resonance forms are the same as the green arrows in the text.

(a)

Move electrons away from (−) charge.

(b)

Move electrons toward (+) charge.

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Organic Chemistry Solutions Manual

Daniel F McAuley

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