

Reactions of Amines and Epoxides

Preparation of Amines

Amines of different types can be prepared by various methods. The following are some of the common methods:



Free electron lone pairs

There are two types of bonding: **sigma bonding** and **covalent bonding**.

- **Sigma bonds** are formed by the overlap of orbitals from one atom to another (homonuclear) (positive) and **covalent** (hetero) (negative).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of σ and π Bonds



Bond and Bonding in π Bonds



Bonding in π Bonds



There are two types of lone pairs

• **Delocalized lone pairs** are found in molecules with conjugated systems.



A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide reactivity and regioselectivity

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is described by the **dipole moment** (μ).



- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).

- δ^- (delta minus) indicates an electron deficient area.
- δ^+ (delta plus) indicates an electron rich area.

Molecular Dipole

To increase the molecular polarity, the same side of the dipole is used.

The polarity of the overall molecule is determined by the vector sum of the dipole moments of the individual bonds.



Linear molecules are more polar than bent molecules.

Average values of C-C and C-H bonds, depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp^3$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp^3$
Length	154 Å	134 Å	113 Å
Strength	376 kJ/mol	473 kJ/mol	368 kJ/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp^3$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp^3$
Length	154 Å	134 Å	113 Å
Strength	410 kJ/mol	463 kJ/mol	439 kJ/mol

Dipole moment values and strength →

Sigma and pi bonds

Covalent bonds are classified as **sigma** and **pi** bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.

- **One single bond** is a sigma bond.
- **Two double bonds** consist of one sigma and one pi bond.



• **pi bonds** are stronger than σ bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be **polar** and **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the electrons are shared equally (nonpolar).
- If the difference in electronegativity is **between 0.5 and 1.7**, the bond is **polar covalent**.
- If the difference in electronegativity is **greater than 1.7**, the bond is **ionic**.

- A difference of **1.7 or higher** is a high threshold for the distinction between ionic and covalent bonds.

Bond Length and Bond Strength

• **Double bond is stronger than single bond.**

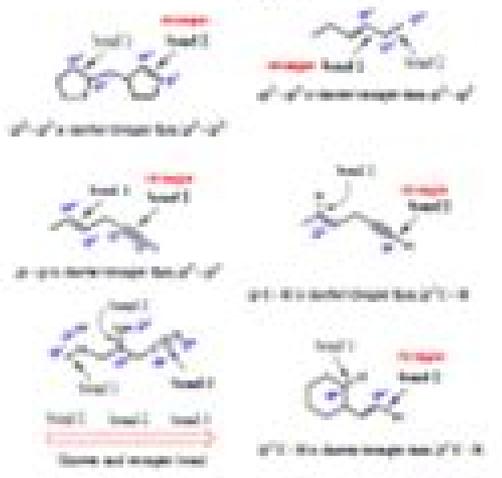
Bond	H-H	H-Cl	H-Br	H-I
Length	74 Å	127 Å	141 Å	154 Å
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds.



• **Double bond is stronger than single bond.**

Example



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