

Reactions of Amines and Epoxides

Preparation of Amines

Amines of different types can be prepared by various methods. The following are some of the common methods:



Free electron lone-pairs

There are two types of bonding: **sigma bonding** and **covalent bonding**.

• **Sigma bonds** are formed by the overlap of orbitals from one atom to another (homonuclear) (positively) and **covalent bonding** (hetero) (stronger).



• **Covalent bonds** are formed by sharing of electrons between two atoms.



There are two main types of difference of orbitals are:



Sigma (σ) and pi (π) bonds

Covalent bonds are classified as **sigma (σ) bonds** and **pi (π) bonds**.

• **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.

• **Pi bonds** are formed by side-on overlap of orbitals.



• Pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be **polar** and **nonpolar**.



• If the difference in electronegativity is **less than 0.5**, the electrons are shared equally (nonpolar). For two atoms, forming a **nonpolar covalent bond**. (For example, C-C and C-H bonds).

• If the difference in electronegativity is **between 0.5 and 1.7**, the electrons are shared unequally (polar). For example, C-O and C-N bonds.

• A difference of **1.7 or higher** is a high difference, and the electrons are not shared (ionic bond).

A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is described by the **dipole moment (μ)**.



• The dipole moment is defined as a vector pointing from the **electropositive** to toward the **more electronegative** atom.

• The more electronegative atom has a **partial negative charge (δ-)** and the less electronegative atom has a **partial positive charge (δ+)**.

• **δ-** (delta minus) indicates an electron-deficient atom.

• **δ+** (delta plus) indicates an electron-rich atom.

Molecular Dipole

To increase the molecular polarity, the same side of the dipole is used.

The polarity of the molecule is a vector sum of the dipole moments of the atoms. A good example is carbon dioxide.



Cancel the vector in the same direction into the vector sum.

Average values of C-C and C-H bonds, depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp^3$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp^3$
Length	153.8	133.4	133.6
Strength	376 kJ/mol	473 kJ/mol	368 kJ/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp^3$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp^3$
Length	109.4	108.4	108.6
Strength	413 kJ/mol	440 kJ/mol	423 kJ/mol

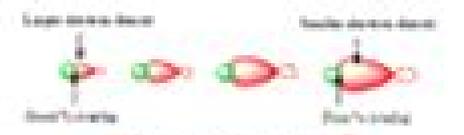
Dipole moment values and strength ↑

Bond Length and Bond Strength

← Bond length increases →

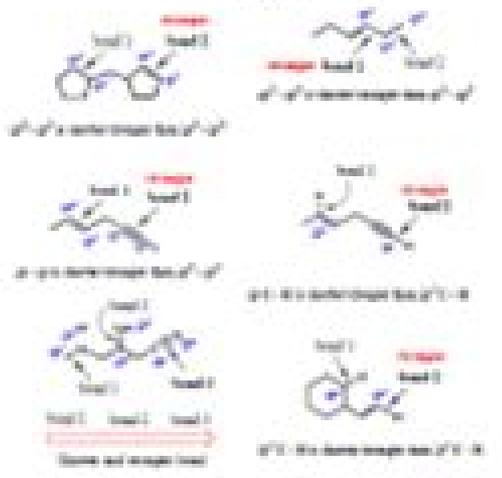
Bond	H-H	H-Cl	H-Br	H-I
Length	74.14	127.4	141.4	154.4
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds, and triple bonds are the strongest.



← Bond strength increases →

Example



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