

Reactions of Amines and Epoxides

Preparation of Amines

Amines of different types can be prepared by various methods. The following are some of the common methods:



Free electron lone-pairs

There are two types of bonding: **sigma bonding** and **covalent bonding**.

- **Sigma bonds** are formed by the overlap of orbitals from one atom to another (homogeneous overlap) and **covalent bonding** (heteroatomic bonding).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of σ and π Bonds



Bond and Bonding in π Bonds



Bonding in π Bonds



There are two types of lone pairs

• **Localized lone pairs**



Sigma and pi (π) bonds

- **Covalent bonds** are classified as **sigma (σ)** and **pi (π)**.
- **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.
- **π bonds** are formed by side-on overlap of orbitals.
- **One single bond** is a **sigma bond**.
- **Two double bonds** consist of **one sigma and one pi bond**.



• π bonds are stronger than σ bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be **polar** and **nonpolar**.



- If the difference in electronegativity is **less than 1.7**, the electrons are shared equally (distribution) for two atoms, forming a **nonpolar covalent bond**. (For example, examples are for C-C and C-H bonds).

- If the difference in electronegativity is **between 1.7 and 1.9**, the bond is **polar covalent bond**. (For example, examples are given for carbon-halogen, N, O, F and other halogens).

- A difference of **1.9 or higher** is so high that the electrons are not shared (bonded) and a **ionic bond** is formed.

A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is described by the **dipole moment (μ)**.

It is a vectorial quantity.



It is the product of the charge and the distance between the charges.

- The dipole moment is directed in a **vector pointing from the less electronegative to toward the more electronegative atom**.
- The more electronegative atom has a **partial negative charge (δ^-)** and the less electronegative atom, a **partial positive charge (δ^+)**.
- **δ^- (delta minus)** indicates an **electron deficient area**.
- **δ^+ (delta plus)** indicates an **electron rich area**.

Molecular Dipole

To increase the molecular polarity, the **vector sum of the dipole moment**.

The **polarity of the molecule is a vectorial sum of the dipole moments of the individual polar bonds** for dipole moment are considered. A good example is carbon dioxide.



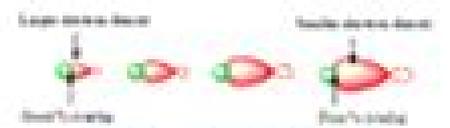
Linear molecule with two polar bonds (C=O) but net dipole moment.

Bond Length and Bond Strength

• **Single bond is the longest and weakest.**

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double and triple bonds are stronger than single bonds and a stronger bond.

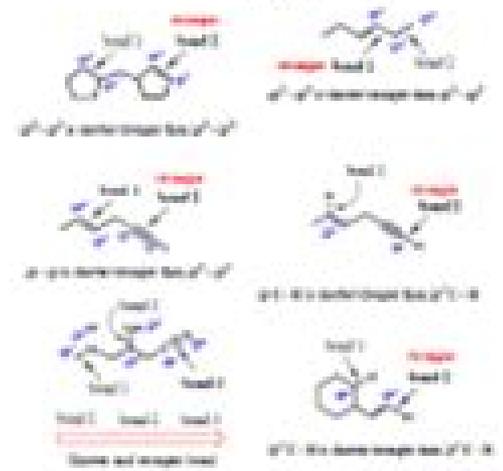


• **Triple bond is the shortest and strongest.**

	Group 1	Group 2	Group 3
C-C Bond	H ₂ C-CH ₃	H ₂ C=CH ₂	H ₂ C≡CH
Substance	ethane	ethylene	acetylene
Length	154 pm	134 pm	120 pm
Strength	376 kJ/mol	611 kJ/mol	837 kJ/mol
Substance	ethane	ethylene	acetylene
Hybridization	sp ³	sp ²	sp
Percentage	25%	33%	50%
Length	154 pm	134 pm	120 pm
Strength	376 kJ/mol	611 kJ/mol	837 kJ/mol

• **Triple bond is the shortest and strongest bond.**

Example



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web you can't stop me you know it's all my fault I'm very very bad and I'm going to make you a vampire John Quinn a made vampire changed by the lamia vampire hunter Redfern himself initially he despised human beings as he considers them being weak beings however everything changes when he met his soulmate

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