

Reactions of Amines and Epoxides

Preparation of Amines

Amines of different types can be prepared by various methods. The following are some of the common methods:



Free electron lone-pairs

There are two types of bonding: **sigma bonding** and **covalent bonding**.

- **Sigma bonds** are formed by the overlap of orbitals from one atom to another (homogeneous overlap) and **covalent bonding** (heteroatomic bonding).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of σ and π Bonds



Bond and Bonding in π Bonds



Bonding in π Bonds



There are two types of π bonds

• **End-to-end overlap**



Sigma and pi (π) bonds

- **Covalent bonds** are classified as **sigma (σ)** and **pi (π)**.
- **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.
- **π bonds** are formed by side-on overlap of orbitals.
- **One single bond** is a **sigma bond**.
- **Two double bonds** consist of **one sigma and one pi bond**.



• π bonds are weaker than σ bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be **polar** and **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the electrons are shared equally (nonpolar bond). For two atoms, forming a **nonpolar covalent bond**. (For example, **C-C** and **C-H** bonds).

- If the difference in electronegativity is **between 0.5 and 1.7**, the electrons are shared unequally (polar covalent bond). (For example, **C-O** and **C-N** bonds).

- A difference of **1.7 or higher** is so high that the electrons are not shared (ionic bond) or **ionic bond** is formed.

A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is described by the **dipole moment (μ)**.

It is a vectorial sum of the electronegativity differences of the two atoms.



Electronegativity differences of the atoms in a bond.

- The dipole moment is shown as a **vector pointing from the less electronegative to the more electronegative atom**.
- The more electronegative atom has a **partial negative charge (δ^-)** and the less electronegative atom has a **partial positive charge (δ^+)**.
- **δ^- (delta minus)** indicates an **electron deficient area**.
- **δ^+ (delta plus)** indicates an **electron rich area**.

Molecular Dipole

To increase the molecular polarity, the **vector sum of the dipoles** is used.

The **polarity of the molecule is a vectorial sum of the dipoles** (if the dipoles are oriented in the same direction).



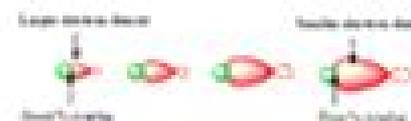
• **Same direction** as the other dipoles (additive effect).

Bond Length and Bond Strength

• **Same bond length, same strength**

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

• **Double bonds are stronger than single bonds** and **stronger than triple bonds**.

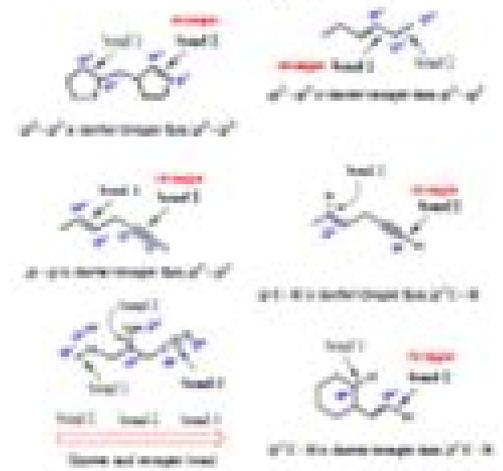


• **Stronger bond, shorter bond length**

Average values of C-C and C-H bonds, depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	473 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	413 kJ/mol	505 kJ/mol	591 kJ/mol

• **Higher s character, shorter and stronger bond**

Example



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L Reisser



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