

RADIOACTIVITY AND RADIOACTIVE DECAY LAB REPORT

I. Abstract

Radioactive elements are substances that go through a process of decay to reach a stable configuration and then emit particles like alpha, beta, and gamma. SI units called Curie (Ci) and Becquerel (Bq) are used to measure radioactivity. Although there are many radioactive elements on Earth, which are constantly decaying to create stable nuclei, these processes could pose major health risks to living things. Consequently, regulation is necessary when operating in vitro. When doing radioactive tests, biosafety limits should be closely adhered to in order to reduce contaminated exposure.

When a parent radionuclide undergoes an alpha, beta, or isomeric transition to become its daughter radionuclide, the decay takes an exponential form, which is indicated by the decay constant λ . The decay constant measures the likelihood that one radioatom will decay for every unit of time. The half-life of a radionuclide—the time at which half of the radioatoms present will have decayed—can be usefully expressed using the decay equation. Because radionuclide half-life influences the capacity to assess tracer kinetics and produce accurate nuclear pictures as well as affects organ, tumor, and whole-body radiation dosage, it has direct consequences for nuclear imaging, radiation therapy, and radiation safety. The mass of radioatoms present in a sample can be calculated to determine the specific activity; the number of radioatoms present in a sample is equal to the activity, defined as the number of transitions per unit time divided by the decay constant. Radioactive equilibrium, which can take one of two forms (secular or transient) and is directly relevant to radionuclides produced by generators, may result from the dynamic relationship between the number of parent and daughter atoms present over time.

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