

16.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide and hydronium ion concentrations from K_w and K_a or K_b and the concentration of a weak acid or base.
3. Calculate the pH of weak acid or base solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion | 5. K_a | 6. K_b |
| | 7. weak base | |

Key Equations

- | | |
|--|---|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ | 4. $\text{pOH} = -\log[\text{OH}^-]$ |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions. Show calculations for all problems.

1. What products are formed when hydrochloric acid reacts with sodium hydroxide? Write the balanced chemical equation for this reaction and identify the products as strong or weak acids/bases.
2. Hydrochloric acid has a concentration of 0.10 M. What is the concentration of H_3O^+ ions in this solution? What is the concentration of OH^- ions in this solution? What is the pH of this solution?
3. What is the pH of a 0.10 M solution of hydrochloric acid? What is the concentration of OH^- ions in this solution? What is the concentration of H_3O^+ ions in this solution?
4. What is the pH of a 0.10 M solution of sodium hydroxide? What is the concentration of H_3O^+ ions in this solution? What is the concentration of OH^- ions in this solution?

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Mike Jess



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