



Unit: Chemical Bonding **Answer Key** "Covalent Bonding" - WS #3

Directions: Please answer each question using your notes from today's class. ☺

1. What is the difference between a molecular formula, structural formula and an electron dot formula? Give an example of each.

A molecular formula gives the chemical make-up of one molecule. It does not show atom-to-atom connectivity. Example: C_2H_2

A structural formula uses lines instead of dots to show shared pairs of electrons.

Example: $H-H$

An electron dot formula uses dots instead of lines to show shared pairs of electrons.

Example: $H:H$

2. What is the difference between single, double, and triple covalent bonds? How many shared electrons are in each?

The difference between the types of bonds is: a single bond is two shared electrons, a double is four shared electrons, and a triple is six shared electrons.

3. Draw the electron dot and structural formulas for the following molecules.

Substance	Rough copy (work!)	Electron dot	Structural
a) Cl_2		<i>Note: The electron dot structures have dots in between each atom instead of lines.</i>	$\text{:}\ddot{\text{Cl}}\text{—}\ddot{\text{Cl}}\text{:}$
b) H_2S			$\begin{array}{c} \text{:}\ddot{\text{S}}\text{:} \\ \\ \text{H—H} \end{array}$
c) BrI			$\text{:}\ddot{\text{I}}\text{—}\ddot{\text{Br}}\text{:}$

Overview Chemical Bonds Answer Key

KJ Lindholm-Leary



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Glencoe Science, 2002 Inorganic Chemistry for JEE Advanced: Part 2, 3E (Free Sample) K. S. Verma, 2022-05-19 Inorganic Chemistry for JEE Advanced Part 2 a Cengage Exam Crack Series product is designed to help aspiring engineers focus on the subject of inorganic chemistry from two standpoints To develop their caliber aptitude and attitude for the engineering field and profession To strengthen their grasp and understanding of the concepts of the subjects of study and their applicability at the grassroots level Each book in this series approaches the subject in a very conceptual and coherent manner While its illustrative solved examples facilitate easy mastering of the concepts and their applications an array of solved problems exposes the students to a variety of questions that they can expect in the examination The coverage and features of this series of books make it highly useful for all those preparing for JEE Main and Advanced and aspiring to become engineers

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Human Biology: Breathing Craig H. Heller, 1999

Chemical Interactions, 2005

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Green Hydrogen Generation from Biomass Yasser Vasseghian, Hesam Kamyab, Changlei Xia, 2026-02-11

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Deciphering the Chemical Code N. D. Epiotis,1996 This groundbreaking work the culmination of more than 10 years of research presents a breakthrough theory of chemical bonding across the periodic table Professor Epiotis an internationally known and respected member of the theoretical community challenges the conventional chemical concepts that underlie popular theories of chemical bonding Building on his insight that electron electron repulsion is the single crucial variable that differentiates one chemical system from another the author formulates explains and applies a new approach based on nonorthogonal valence bond methodology that amounts to nothing less than a revolutionary unified theory of chemical bonding across the periodic table This work represents the first post Pauling theory of chemical bonding New theory means new formulae and this work is about new chemical formulae that lead to the self consistent rationalization of existing facts and even more important the design of new chemistry

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