

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~Sulfuric acid~~ **H<sub>2</sub>SO<sub>4</sub>**
- ~~potassium hydroxide~~ **KOH**
- ~~Sodium hydroxide~~ **NaOH**
- ~~calcium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~ammonium hydroxide~~ **NH<sub>4</sub>OH**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.  
HClO<sub>4</sub> is STRONG with a lower cell completely dissociate in water

- HCl**      ~~HCl~~      ~~HNO<sub>3</sub>~~      ~~Mg(OH)<sub>2</sub>~~      **HNO<sub>3</sub>**      **H<sub>2</sub>SO<sub>4</sub>**      ~~HNO<sub>2</sub>~~      **HClO<sub>4</sub>**

3. What is the formula for hydronium? **H<sub>3</sub>O<sup>+</sup>**

4. H<sub>2</sub>CO<sub>3</sub> will react with water and lose a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction: **H<sub>2</sub>CO<sub>3</sub> + H<sub>2</sub>O ⇌ H<sub>3</sub>O<sup>+</sup> + HCO<sub>3</sub><sup>-</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H<sub>2</sub>CO<sub>3</sub> is a weak acid.**
- Would you define H<sub>2</sub>CO<sub>3</sub> as a proton donor or a proton acceptor? **proton donor, because it is losing a H<sup>+</sup> ion.**
- In your written equation, is H<sub>2</sub>CO<sub>3</sub> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of H<sub>2</sub>CO<sub>3</sub>: **HCO<sub>3</sub><sup>-</sup>**
- Identify H<sub>2</sub>CO<sub>3</sub> as polyprotic, amphoteric, or neither: **polyprotic, because it contains multiple H's to give.**

5. F<sup>-</sup> will react with water and gain a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction: **F<sup>-</sup> + H<sub>2</sub>O ⇌ HF + OH<sup>-</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F<sup>-</sup> is a weak base.**
- Would you define F<sup>-</sup> as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H<sup>+</sup> ion.**
- In your written equation, is F<sup>-</sup> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is getting a proton.**
- Identify the conjugate acid of F<sup>-</sup>: **HF**
- Identify F<sup>-</sup> as polyprotic, amphoteric, or neither: **neither.**

6. HSO<sub>4</sub><sup>-</sup> will react with water and lose a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction: **HSO<sub>4</sub><sup>-</sup> + H<sub>2</sub>O ⇌ H<sub>3</sub>O<sup>+</sup> + SO<sub>4</sub><sup>2-</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO<sub>4</sub><sup>-</sup> is a weak acid.**
- Would you define HSO<sub>4</sub><sup>-</sup> as a proton donor or a proton acceptor? **proton donor, because it is losing a H<sup>+</sup> ion.**
- In your written equation, is HSO<sub>4</sub><sup>-</sup> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of HSO<sub>4</sub><sup>-</sup>: **SO<sub>4</sub><sup>2-</sup>**
- Identify HSO<sub>4</sub><sup>-</sup> as polyprotic, amphoteric, or neither: **amphoteric, because of both a negative charge (only will take a proton) and a H to give (can give out H's too).**

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