



CHEMISTRY

CHEAT SHEET



NAMING SIMPLE COMPOUNDS

NAMING SIMPLE COMPOUNDS CHEAT SHEET

Reading: Compounds with their formulas
from the ions in the periodic table

IONIC

Reading: Naming Compounds

Directions for naming:

- Show the cation first.
- Show the anion second.
- Use the full name of the cation.
- Use the full name of the anion.
- Use the charge of the cation and the charge of the anion.
- Use the charge of the cation and the charge of the anion.
- Use the charge of the cation and the charge of the anion.
- Use the charge of the cation and the charge of the anion.

Cations		Anions	
Symbol	Name	Symbol	Name
Li ⁺	Lithium	F ⁻	Fluoride
Na ⁺	Sodium	Cl ⁻	Chloride
K ⁺	Potassium	Br ⁻	Bromide
Rb ⁺	Rubidium	I ⁻	Iodide
Cs ⁺	Cesium	S ²⁻	Sulfide
Ba ²⁺	Barium	O ²⁻	Oxide
Ca ²⁺	Calcium	Se ²⁻	Selenide
Mg ²⁺	Magnesium	Te ²⁻	Telluride
Zn ²⁺	Zinc	P ³⁻	Phosphide
Fe ²⁺	Iron (II)	As ³⁻	Arsenide
Fe ³⁺	Iron (III)	Sb ³⁻	Antimonide
Co ²⁺	Cobalt (II)	Bi ³⁻	Bismuthide
Ni ²⁺	Nickel (II)	Sn ²⁺	Stannous
Cu ⁺	Copper (I)	Sn ⁴⁺	Stannic
Cu ²⁺	Copper (II)	Pb ²⁺	Plumbous
Ag ⁺	Silver	Pb ⁴⁺	Plumbic
Au ⁺	Gold (I)		
Au ³⁺	Gold (III)		

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Reading: Compounds with a central atom

MOLECULAR

Directions for naming:

- Show the central atom first.
- Show the number of atoms of the central atom.
- Show the number of atoms of the other atom.
- Show the name of the other atom.
- Show the charge of the central atom and the charge of the other atom.
- Show the charge of the central atom and the charge of the other atom.
- Show the charge of the central atom and the charge of the other atom.
- Show the charge of the central atom and the charge of the other atom.

Reading: Names

Formula	Name	Formula	Name
CO ₂	Carbon dioxide	SO ₂	Sulfur dioxide
CO	Carbon monoxide	SO ₃	Sulfur trioxide
NO ₂	Nitrogen dioxide	NO	Nitrogen monoxide
NO	Nitrogen monoxide	NO ₂	Nitrogen dioxide
Cl ₂	Chlorine	Cl ₂ O	Dichlorine monoxide
O ₂	Oxygen	Cl ₂ O ₇	Dichlorine heptoxide
H ₂	Hydrogen		
N ₂	Nitrogen		
Br ₂	Bromine		
I ₂	Iodine		
S ₈	Sulfur		
P ₄	Phosphorus		
As ₄	Arsenic		
Sb ₃	Antimony		
Bi ₃	Bismuth		

Common Names for Specific Compounds

Formula	Common Name	Formula	Common Name
NaCl	Salt	NaOH	Sodium hydroxide
CaCl ₂	Calcium chloride	Ca(OH) ₂	Slaked lime
CaSO ₄	Plaster of Paris	CaCO ₃	Limestone
CaH ₂	Calcium hydride	CaO	Quicklime
CaF ₂	Fluorite		

Naming Compounds Study Guide

D Keegan



Naming Compounds Study Guide:

Unveiling the Energy of Verbal Artistry: An Mental Sojourn through **Naming Compounds Study Guide**

In a world inundated with screens and the cacophony of instant transmission, the profound energy and mental resonance of verbal artistry often diminish in to obscurity, eclipsed by the constant barrage of sound and distractions. However, nestled within the lyrical pages of **Naming Compounds Study Guide**, a captivating perform of fictional elegance that impulses with raw thoughts, lies an remarkable trip waiting to be embarked upon. Published by way of a virtuoso wordsmith, this interesting opus courses visitors on a psychological odyssey, delicately revealing the latent possible and profound impact embedded within the complicated internet of language. Within the heart-wrenching expanse with this evocative analysis, we shall embark upon an introspective exploration of the book is central subjects, dissect its fascinating writing fashion, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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