

Advance Study Assignment: pH, Its Measurement and Applications

1. A solution of a weak acid was tested with the indicators used in this experiment. The colors observed were as follows:

Methyl violet	violet	Congo red	violet
Thymol blue	orange	Bromcresol green	yellow
Methyl yellow	red		

What is the approximate pH of the solution?

2. An aqueous solution of NH_3 has a pH of 11.6. The ammonia (NH_3) molecule is the conjugate base of the NH_4^+ (called "ammonium") ion. Write the net ionic equation for the reaction that makes an aqueous solution of NH_3 basic. (Eq. 5.)

3. The pH of a 0.10 M HOBr solution is 4.8.

a. What is $[\text{H}^+]$ in that solution?

_____ M

b. What is $[\text{OBr}^-]$? What is $[\text{HOBr}]$? (Where do the H^+ and OBr^- ions come from?)

_____ M; _____ M

c. What is the value of K_a for HOBr? What is the value of $\text{p}K_a$?

4. Formic acid, HFor, has a K_a value equal to about 1.8×10^{-4} . A student is asked to prepare a buffer having a pH of 3.90 from a solution of formic acid and a solution of sodium formate having the same molarity. How many milliliters of the NaFor solution should she add to 20. mL of the HFor solution to make the buffer? (See discussion of buffers.)

_____ mL

5. How many mL of 0.10 M NaOH should the student add to 20 mL 0.10 M HFor if she wished to prepare a buffer with a pH of 3.90, the same as in Problem 4?

_____ mL

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