

## ELECTRIC CHARGE AND ELECTRIC FIELD

- 21.1.** **(a) IDENTIFY and SETUP:** Use the charge of one electron ( $-1.602 \times 10^{-19}$  C) to find the number of electrons required to produce the net charge.  
**EXECUTE:** The number of excess electrons needed to produce net charge  $q$  is  

$$\frac{q}{-e} = \frac{-3.20 \times 10^{-9} \text{ C}}{-1.602 \times 10^{-19} \text{ C/electron}} = 2.00 \times 10^{10} \text{ electrons.}$$
**(b) IDENTIFY and SETUP:** Use the atomic mass of lead to find the number of lead atoms in  $8.00 \times 10^{-3}$  kg of lead. From this and the total number of excess electrons, find the number of excess electrons per lead atom.  
**EXECUTE:** The atomic mass of lead is  $207 \times 10^{-3}$  kg/mol, so the number of moles in  $8.00 \times 10^{-3}$  kg is  

$$n = \frac{m_{\text{lead}}}{M} = \frac{8.00 \times 10^{-3} \text{ kg}}{207 \times 10^{-3} \text{ kg/mol}} = 0.03865 \text{ mol.}$$
 $N_A$  (Avogadro's number) is the number of atoms in 1 mole, so the number of lead atoms is  $N = nN_A = (0.03865 \text{ mol})(6.022 \times 10^{23} \text{ atoms/mol}) = 2.328 \times 10^{22}$  atoms.  
The number of excess electrons per lead atom is  $\frac{2.00 \times 10^{10} \text{ electrons}}{2.328 \times 10^{22} \text{ atoms}} = 8.59 \times 10^{-13}$ .  
**EVALUATE:** Even this small net charge corresponds to a large number of excess electrons. But the number of atoms in the sphere is much larger still, so the number of excess electrons per lead atom is very small.
- 21.2.** **IDENTIFY:** The charge that flows is the rate of charge flow times the duration of the time interval.  
**SETUP:** The charge of one electron has magnitude  $e = 1.60 \times 10^{-19}$  C.  
**EXECUTE:** The rate of charge flow is 20,000 C/s and  $t = 100 \mu\text{s} = 1.00 \times 10^{-4}$  s.  

$$Q = (20,000 \text{ C/s})(1.00 \times 10^{-4} \text{ s}) = 2.00 \text{ C.}$$
The number of electrons is  $n_e = \frac{Q}{1.60 \times 10^{-19} \text{ C}} = 1.25 \times 10^{20}$ .
- 21.3.** **IDENTIFY:** From your mass estimate the number of protons in your body. You have an equal number of electrons.  
**SETUP:** Assume a body mass of 70 kg. The charge of one electron is  $-1.60 \times 10^{-19}$  C.  
**EXECUTE:** The mass is primarily protons and neutrons of  $m = 1.67 \times 10^{-27}$  kg. The total number of protons and neutrons is  $n_{p+ne} = \frac{70 \text{ kg}}{1.67 \times 10^{-27} \text{ kg}} = 4.2 \times 10^{28}$ . About one-half are protons, so  
 $n_p = 2.1 \times 10^{28} = n_e$ . The number of electrons is about  $2.1 \times 10^{28}$ . The total charge of these electrons is  

$$Q = (-1.60 \times 10^{-19} \text{ C/electron})(2.1 \times 10^{28} \text{ electrons}) = -3.35 \times 10^9 \text{ C.}$$
**EVALUATE:** This is a huge amount of negative charge. But your body contains an equal number of protons and your net charge is zero. If you carry a net charge, the number of excess or missing electrons is a very small fraction of the total number of electrons in your body.

# Physics Electric Field Study Guide

**John Frank Friese**



## **Physics Electric Field Study Guide:**

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### **Table of Contents Physics Electric Field Study Guide**

1. Understanding the eBook Physics Electric Field Study Guide
  - The Rise of Digital Reading Physics Electric Field Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Physics Electric Field Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physics Electric Field Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Electric Field Study Guide
  - Personalized Recommendations
  - Physics Electric Field Study Guide User Reviews and Ratings
  - Physics Electric Field Study Guide and Bestseller Lists

5. Accessing Physics Electric Field Study Guide Free and Paid eBooks
  - Physics Electric Field Study Guide Public Domain eBooks
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6. Navigating Physics Electric Field Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Physics Electric Field Study Guide Compatibility with Devices
  - Physics Electric Field Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physics Electric Field Study Guide
  - Highlighting and Note-Taking Physics Electric Field Study Guide
  - Interactive Elements Physics Electric Field Study Guide
8. Staying Engaged with Physics Electric Field Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physics Electric Field Study Guide
9. Balancing eBooks and Physical Books Physics Electric Field Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physics Electric Field Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Physics Electric Field Study Guide
  - Setting Reading Goals Physics Electric Field Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physics Electric Field Study Guide
  - Fact-Checking eBook Content of Physics Electric Field Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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