

Boyle's Law

- 1. The product of a fixed amount of gas varies inversely with the pressure of constant temperature.
- 2. Boyle's
- 3. Law

Kinetic Theory and Boyle's Law

- 1. Pressure of a gas is caused by the collisions of the gas molecules themselves and the container.
- 2. If the gas is compressed to half the volume it has twice as many collisions with the pressure increases twice.
- 3. Force is many times greater per area on the walls of the container.
- 4. Pressure of the gas will double.

Ex 1. A certain quantity of gas has a volume of 200 ml at standard atmospheric pressure. After the volume is reduced, it has a pressure of 1.25 atm above the pressure inside the tube. What is the volume of the gas at this pressure?

$$P_1 V_1 = P_2 V_2 \quad \text{or} \quad 1 \text{ atm} \times 200 \text{ ml} = 1.25 \text{ atm} \times V_2$$

$$V_2 = \frac{200 \text{ ml}}{1.25} = 160 \text{ ml}$$

Ex 2. A gas is pressure of 1.25 atm and has a volume of 200 ml. If the pressure is increased until the volume is 160 ml, what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad \text{or} \quad 1.25 \text{ atm} \times 200 \text{ ml} = P_2 \times 160 \text{ ml}$$

$$P_2 = \frac{1.25 \text{ atm} \times 200 \text{ ml}}{160 \text{ ml}}$$

$$P_2 = 1.5625 \text{ atm}$$

Physics Episode 902 Note Taking Guide Answers

Guillaume Favre



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